**Reaction A**

1. Pour 50 mL of the copper (II) chloride solution into the 100 mL beaker.
2. Loosely crumple the piece of aluminum foil.
3. Drop it into the 100 mL beaker.
4. When the reaction stops, rinse contents of the beaker down the sink with a **large** amount of water.

**Reaction B**

1. Add 10 mL of 6.0 M sulfuric acid to one of the test tubes.
2. Obtain two pieces of zinc from the chemical area.
3. Hold the test tube with the metal holder.
4. Tilt the test tube at an approximately 45 degree angle.
5. Light one of the wooden splints.
6. Carefully slide one of the pieces of zinc into the test tube.
7. After 15 seconds, hold the lit wooden splint near the mouth of the test tube.
8. When the reaction stops, pour the contents of the test tube into the **labeled waste** **container** on the chemical table.

**Reaction C**

1. Pour 10 mL of potassium iodide into one of the test tubes.
2. RINSE the graduated cylinder well.
3. Pour 10 mL of lead (II) nitrate into the same test tube.
4. When the reaction settles, pour the contents into the sink and flush with **a lot** of water.

**Reaction D**

1. Put on a pair of green chemical gloves.
2. Measure out 50 mL of concentrated peroxide and pour into the 100mL cylinder.
3. Add one squirt of liquid dish detergent.
4. Return to your lab station and swirl the graduated cylinder OVER THE BIN until the two substances completely mix.
5. At the chemical table, pour out approximately 10mL of potassium iodide.
6. Pour the 10mL into the graduated cylinder containing the peroxide and soap.
7. Light a wooden splint, gently blow it out, and immediately touch the bubbles.
8. Relight the splint and hold it against a bubble.

**Clean Up\*\*\*\*\*\*\*\*\*\*\*\*\*\*\***

**Use the bristle brush and soap to scrub out the beakers, the test tubes and the small graduated cylinders. Rinse out the 100mL graduated cylinder and plastic bin until they are completely clean.**